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Communication

Gel electrolytes of covalent network polybenzimidazole and phosphoric acid by direct casting^a

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Polybenzimidazole membranes imbibed with phosphoric acid can support high proton conductivity at 120-200 °C, and have therefore emerged as the state-of-the-art electrolytes for fuel cells operating in this temperature range. This work presents a novel and operationally simple methodology for preparing mechanically robust covalent network polybenzimidazole membranes containing up to 95 wt% phosphoric acid. Diamino-terminal pre-polymers of different chain lengths were first prepared, followed by addition of a trifunctional carboxylic acid. The crude solutions were cast and subsequently heat treated at up to 230 °C, yielding free-standing membranes of networked polybenzimidazole with high proton conductivity at up to 180 °C and encouraging fuel cell performance.

^a **Supporting Information** is available online from the Wiley Online Library or from the author.

1. Introduction

Binary systems of phosphoric acid immobilized in polymer matrices can support high proton conductivity at low water activities up to 200 °C, and are thus widely employed as high temperature polymer electrolyte membranes.^[1] The concept was first proposed by Wainright et al.,^[2] who replaced water with phosphoric acid as the proton solvent in perfluorosulfonic acid membranes. However, due to severe complications with concentration polarization the practical value as fuel cell electrolyte was limited.^[3] Immobilizing phosphoric acid in poly[2,2'-(*m*-phenylene)-5,5'-bibenzimidazole] (*m*-PBI, see structure in Scheme 1) was a considerably more successful approach and its discovery resulted in a major breakthrough in the field.^[4]

A large number of polybenzimidazole structure derivatives,^[5] polyimides^[6] and aromatic polysulfones bearing pyridine units in the main chain^[7] or as tethered side-groups^[8] have been developed as base materials for this purpose throughout the years. The chemistry of such semi-solid polymer electrolytes is highly complex as they are in addition to the base polymer composed of water as well as various phosphoric oxo-acid species.^[9] Although the physicochemical properties depend on the preparation method,^[10] a general understanding is that the proton conductivity at a certain temperature and water activity increases with increasing acid content. For example, for *m*-PBI with an acid content of 4-10 H₃PO₄ per polymer repeat unit it typically reaches 0.04-0.08 S cm⁻¹ at 180 °C.^[11] Under similar conditions the conductivity of alternative polybenzimidazole structures prepared according to the sol-gel method with >20 H₃PO₄ per polymer repeat can reach as high as 0.2 S cm⁻¹.^[12]

The phosphoric acid within the structure indeed supports remarkably high proton conductivity, but it also exerts a strong plasticizing effect. As a result, the tensile strength^[13] as well as the creep resistance^[14] decreases and improving the mechanical robustness at high acid loadings has therefore been a key strategy in the development process. Here, increasing the linear molecular

weight of the polymer,^[15] reinforcement with porous polytetrafluoroethylene^[16] or nanoparticles,^[17] and various crosslinking methods involving thermal treatment,^[18] Friedel-Crafts chemistry,^[19] *N*-functionalization,^[20] phosphoric acid ester bridging^[21] or electron beam irradiation^[22] have proven successful. Addition of trifunctional carboxylic acid monomers to *N*-terminal pre-polymers represent an interesting approach in connection to covalent crosslinking, but the obtained materials present severe processing difficulties due to the formation of an insoluble covalent network polymer.^[23]

In this work, the sol-gel processing methodology^[12, 24] was combined with the covalent network chemistry as described by Xu et al.^[23a, 23b] This was achieved by synthesizing *m*-PBI pre-polymers of different chain lengths and with diamino-functional terminal groups, followed by addition of a trifunctional carboxylic acid before tape-casting on a glass substrate. A subsequent heat treatment produced gel membranes with high phosphoric acid content, showing increased tensile strength and conductivity well above 100 mS cm⁻¹ at 120-180 °C.

2. Experimental Section

2.1. Materials

3,3'-diaminobenzidine (DAB, ≥ 98%) was acquired from Glentham Life Sciences, stored at -18 °C and dried at 110 °C under vacuum overnight prior to use. Isophthalic acid (IPA, 99%, dried under vacuum at 110 °C), benzene-1,3,5-tricarboxylic acid (BTA, 95%, Sigma-Aldrich), orthophosphoric acid (PA, 85% H₃PO₄), polyphosphoric acid (PPA, 115% PA, reagent grade), LiCl (technical grade), potassium hydrogen phthalate (PHP), sodium hydroxide (NaOH), *N,N*-dimethylacetamide (DMAc), dimethylsulfoxide (DMSO) were supplied by Sigma-Aldrich and used as received unless otherwise stated.

2.2. Polymerization and film formation

The pre-polymers were obtained by mixing DAB and IPA in PPA in a reaction vessel equipped with an overhead mechanical stirrer, argon in- and outlet and thermometer. The stoichiometry was controlled to give diamino-terminal pre-polymers with theoretical degrees of polymerization of 25, 50 and 75 (hereafter referred to as DP25, DP50 and DP75, respectively), see Table S1 for details. As a general procedure, the mixture was heated to 150 °C under a flow of argon and stirred for 20 h to form a homogeneous solution. The temperature was thereafter increased to 220 °C and the homogeneous solution was stirred for additionally 4 h and thereafter cooled to 120 °C. A sample was collected, acid-neutralized and used for further analysis. BTA was added to the remaining solution, which was stirred at 120 °C for 30 minutes before the crude solution was cast to a thickness of 300 μm on a glass substrate using a doctorblade. The produced film was then heat-treated at 150 °C for 12 h and at 230 °C for 4 h. Heat-treatment was performed in air using a heating rate of 60 °C h⁻¹. To slowly hydrolyze the PPA to PA, the films were placed in a closed container over an aqueous solution of LiCl (1 wt%) for at least 72 h to adjust the nominal relative humidity to around 88%. The obtained membranes were subsequently equilibrated at room temperature in 85 wt% H₃PO₄ and stored in the acid until use.

2.3. Characterization

The gel content of the membranes was determined on the dry membrane basis (after extensive washing in water and dilute aqueous sodium hydroxide (hereafter referred to as de-doping) by treating the samples in DMAc at 120 °C for 24 h. The PA content was determined by extracting 4 × 1 cm² samples in 400 mL water for 24 h, which was titrated with a standardized aqueous NaOH solution using an 808 Titrando from Metrohm. The PA content was calculated based on the second equivalence point. For comparison, the PA content was also determined gravimetrically based on the weight before and after de-doping. The viscosity was determined at 30 °C in DMSO or 96 % H₂SO₄

at a sample concentration of about 0.5 g dL⁻¹ using an Ubbelohde viscometer. The Fourier transform infrared spectra were recorded on a Perkin Elmer Spectrum Two in attenuated total reflectance (ATR) mode. The ¹H nuclear magnetic resonance (¹H NMR) spectra were recorded on a Bruker Ascend operating at 400 MHz, using deuterated dimethylsulfoxide (DMSO-*d*₆) as solvent and using the solvent residual signal at 2.50 ppm as internal reference. Thermogravimetric data were acquired under air at a linear heating rate of 5 °C min⁻¹, using a Netsch STA 409 PC. X-ray diffraction (XRD) was carried out using a Rigaku MiniFlex 600 equipped with a Cu K_α X-ray source ($\lambda = 1.5418 \text{ \AA}$) in the range of 5-90° 2 θ at a scan rate of 2° 2 θ min⁻¹ and a step size of 0.02° 2 θ . Tensile testing was performed using a modified universal materials testing machine (Testometric Micro 350) at a crosshead speed of 10.00 mm min⁻¹. Dog-bone shaped specimens (2.0 mm wide and 28 mm long) were prepared by die-cutting and dried in vacuum at 60 °C for at least 12 hours. The in-plane conductivity was determined at 120-180 °C under ambient air using a four-probe in-plane conductivity cell. A symmetric square wave current (2-5 kHz) was applied via platinum foils and the voltage drop across 1 cm was measured using an Yokogawa DL1100 oscilloscope.

Fuel cell membrane-electrode assemblies were prepared by sandwiching the DP50 membrane between two pieces of HT140E-W gas-diffusion electrodes with a Pt loading of 0.5 mg cm⁻² acquired from E-Tek (active area 1 cm²). The cells were operated by passing air at a rate of 150 mL min⁻¹ over the cathode and dry hydrogen at a rate of 50 mL min⁻¹ over the anode at ambient pressure. Constant current operation was performed at 200 mA cm⁻².

3. Results and Discussion

By controlling the relative DAB:IPA molar ratio during the pre-polycondensation at 1.040-1.013, linear diamino-terminal pre-polymers were obtained with targeted degrees of polymerization of 25,

50 and 75. The identity of the linear pre-polymers was confirmed by ^1H NMR and FT-IR, showing the main spectral features of *m*-PBI as assigned and discussed in more detail elsewhere.^[25]

The ^1H NMR of the DP25 pre-polymer is shown in Figure 1a, showing several peaks in the region 6.5-7.0 ppm. This is a characteristic region for aromatic protons of the amino-terminal groups (compare with the spectrum of the DAB monomer), and supports that the pre-polymers contain active terminal groups for further network formation. Furthermore, as shown in Figure 1b the FT-IR of the DP25 pre-polymer showed a strong broad absorption peak around 3000 cm^{-1} , which is characteristic for hydrogen-bonded N-H groups.^[26] As shown in Figure 1c, the thermo-oxidative stability increased with increasing chain length of the linear pre-polymer, showing onset temperatures of major decomposition of 500, 513 and $570\text{ }^\circ\text{C}$ for DP25, DP50 and DP75, respectively.

The inherent viscosity of the de-doped polymer sampled from the crude solution after the pre-polymerization increased with increasing theoretical degree of polymerization, as shown in Figure 1d. For example, the inherent viscosity of DP25 as measured in H_2SO_4 was $0.51\text{-}0.81\text{ dL g}^{-1}$, corresponding to about 28000 g mol^{-1} .^[27] For the DP50 and DP75 samples, the inherent viscosity was as high as $1.11\text{-}1.55$ and $1.92\text{-}3.10\text{ dL g}^{-1}$, respectively. Significant batch-to-batch deviations were observed as a result of small variation in the reaction conditions.^[28] The inherent viscosity of the linear *m*-PBI pre-polymers was considerably higher than what is normally observed for *m*-PBI in sulfuric acid,^[15b] likely due to altered chain conformation in solution due to the charged terminal di-amino groups.

The process for obtaining the free-standing membranes based on a *m*-PBI network structure is illustrated in Scheme 1. The degree of polymerization of the linear pre-polymers determined the spacing between three-way connections in the network structure of the polymer.

As shown in the inset in Scheme 1, the cast crude solution appeared as a transparent yellow film. The film contracted during the heat treatment resulting in decreased surface area and increased

thickness. However, dark brown free-standing films were obtained after hydrolysis of the PPA in air at room temperature and 88% relative humidity for 72 h. The gel content of the polymer fraction of the de-doped membranes, i.e. the fraction of the polymer that is insoluble in DMAc, was 84, 75 and 72 wt% for DP25, DP50, and DP75, respectively (see Figure S1). This qualitative measure can be taken as an indicator of the degree of crosslinking,^[18, 20a] and the high gel contents of the membranes in the present work clearly support the polymer network formation.

To prevent further leaching of PA, the membranes were equilibrated and stored in 85% PA. The phosphoric acid contents are summarized in Table 1, and ranged of about 91-95 wt%. This corresponds to 39-55 H₃PO₄ per polymer repeat unit assuming 308 g mol⁻¹ as the repeat unit molar mass, similar to that obtained for direct cast membranes based on linear polybenzimidazole chemistries.^[12, 14]

As shown in Figure 2a, the XRD of the membrane after the heat treatment step showed a broad peak centered around 2θ 25° for DP50 and DP75, as commonly observed for phosphoric acid doped *m*-PBI membranes and indicating an amorphous structure with some degree of order on the sub-nanometer scale after heat treatment.^[10, 18] The amorphous halo of the DP25 membrane was centered at 2θ 18°, likely due to the relatively short linear polymer segments. After equilibration in phosphoric acid, the membranes remained fully amorphous.

Due to the high acid-contents, the network membranes showed significantly lower tensile strength and elastic modulus as compared with the reference membrane prepared by casting from DMAc (Figure 2b). The tensile strength at break (around 1 MPa) was, however, similar to that of membranes based on linear *m*-PBI with lower acid contents as reported in the literature.^[10] The elastic modulus of the DP25 membrane with an acid content corresponding to 49-55 H₃PO₄ per polymer repeat unit was about 6.6 MPa, which can be compared with 1.9 for a membrane based on linear *p*-PBI with an acid content of 28 H₃PO₄ per polymer repeat unit.^[29] The relatively high elastic modulus

of the DP25 membrane, as commonly observed for covalently crosslinked *m*-PBI,^[20b] further supports the polymer network formation. Despite slightly lower acid contents, the elastic modulus decreased with increasing chain length of the pre-polymers, as summarized in Table 1. At the same time, the elongation at break increased due the increasing spacing between the three-way connections in polymer network structure.

As shown in Figure 3a, the proton conductivity at 120 °C was initially very high and reached around 0.3 S cm⁻¹, i.e. close to that of pure orthophosphoric acid^[30] and in good agreement with the data reported for direct cast membranes based on linear polybenzimidazole with similar acid contents.^[12, 21] However, due to loss of water during the initial hours^[28] the conductivity dropped significantly and stabilized at around 0.1 S cm⁻¹, which is still higher than that of the membrane cast from DMAc with an acid content of about 10 H₃PO₄ per polymer repeat unit. The conductivity slightly increased by further heating to 160 °C, and thereafter declined slightly at 180 °C due to phosphoric acid condensation.^[18, 31] A deeper quantitative discussion about the sample-to-sample conductivity variations for the prepared membranes is likely not feasible at this stage, due to e.g. uncertainties in compositional and dimensional changes of the membranes under the experimental conditions.

Fuel cell polarization curves of a MEA based on the DP50 membrane are shown in Figure 3b, together with polarization curves of reference cells based on post-doped *m*-PBI cast from DMAc. The cells were operated at high over-stoichiometry of both hydrogen and air (>30), as an initial assessment of the new electrolyte. The higher activation loss as observed in the low current density range of the DP50 cell is likely due to partial phosphoric acid flooding of the catalyst layers.^[32] It should also be remarked that the catalyst loading of the electrodes was 3 times lower than that of the reference cell, which has a large impact on the activation overpotential. In the higher current density range, the electrode flooding is evidenced by mass transport limitations, particularly at 120 and 150 °C.

Furthermore, the DP50 membrane was considerably thicker than the linear reference membrane. In order to compensate for the thickness differences, the in-situ conductivities were calculated based on an estimation of the ohmic resistance by a linear fit to the polarization data in the current density range 350-1170 mA cm⁻². For the DP50 membrane, the in-situ conductivity reached 106 mS cm⁻¹ and was in good agreement with the data in Figure 3a. Even though the area specific resistance of the cell based on the linear reference membrane was comparable to that of the DP50, the corresponding in-situ conductivity was considerably lower (21 mS cm⁻¹).

As stated above, the high phosphoric acid loadings present major challenges with respect to performance optimization. The phosphoric acid tend to redistribute during cell operation,^[33] and further optimization of the electrodes and of the membrane-electrode interfaces is thus needed to maintain the phosphoric acid within the novel networked membranes and ultimately improve their cell performance. Acid leaching and redistribution represent predominating fuel cell degradation mechanisms,^[34] and mechanically robust membranes with exceptionally high acid inventory could therefore become an important part of the overall degradation mitigation strategy. As shown in the inset in Figure 3b, the preliminary short-term stability tests indicate stable performance during a few days of operation.

4. Conclusions

Addition of a trifunctional carboxylic acid to the crude phosphoric acid solution of diamino-terminal polybenzimidazole pre-polymers produced networked polybenzimidazole membranes after casting and heat treatment. The method produced mechanically robust membranes with phosphoric acid contents as high as 90-95 wt%, supporting proton conductivity of around 0.1 S cm⁻¹ at 120-180 °C. The implementation of high phosphoric acid content membranes in fuel cells could potentially

contribute to better performance and longer lifetimes, but introduces challenges with respect to optimization of electrodes and membrane-electrode interfaces.

Supporting Information

Supporting Information is available from the Wiley Online Library or from the author.

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Keywords: polybenzimidazole, network, direct casting, polymer electrolyte, fuel cell

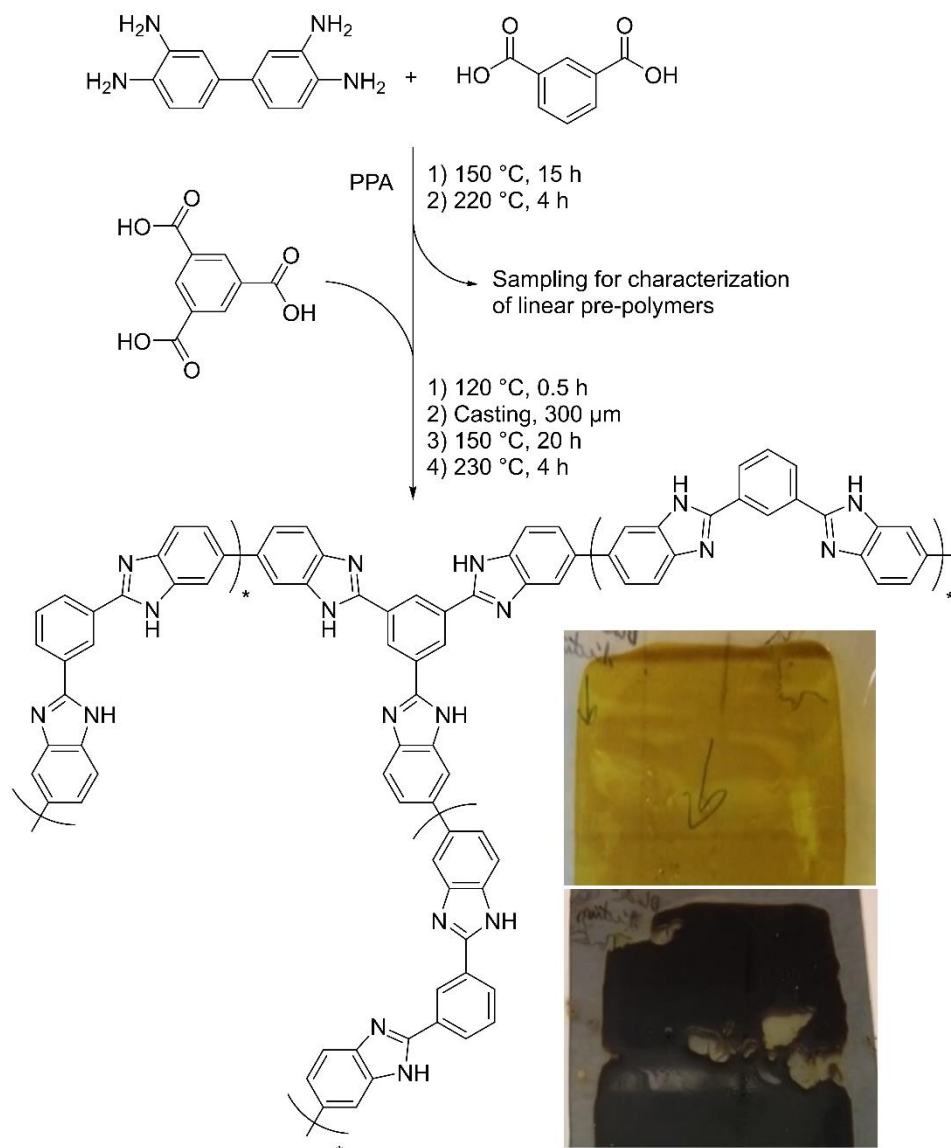
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Scheme 1. Reaction scheme of the linear pre-polymer preparation and the subsequent network formation. Photographs of the film obtained for DP50 after casting of the crude solution (top) and after the subsequent heat treatment (bottom).

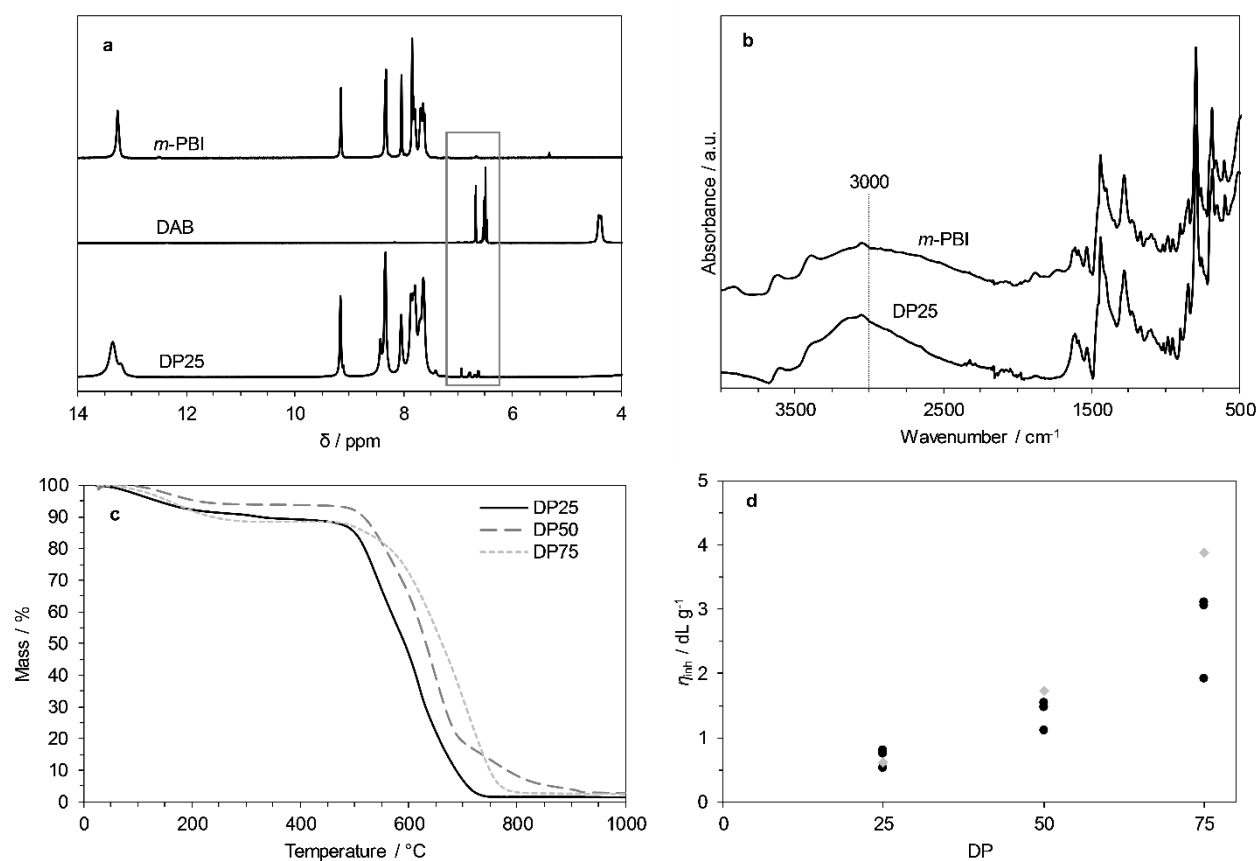


Figure 1. ^1H NMR (a), FTIR (b) and TGA (c) of the linear *m*-PBI pre-polymer and the corresponding inherent viscosities (d) as measured in H_2SO_4 (black) or DMSO (gray).

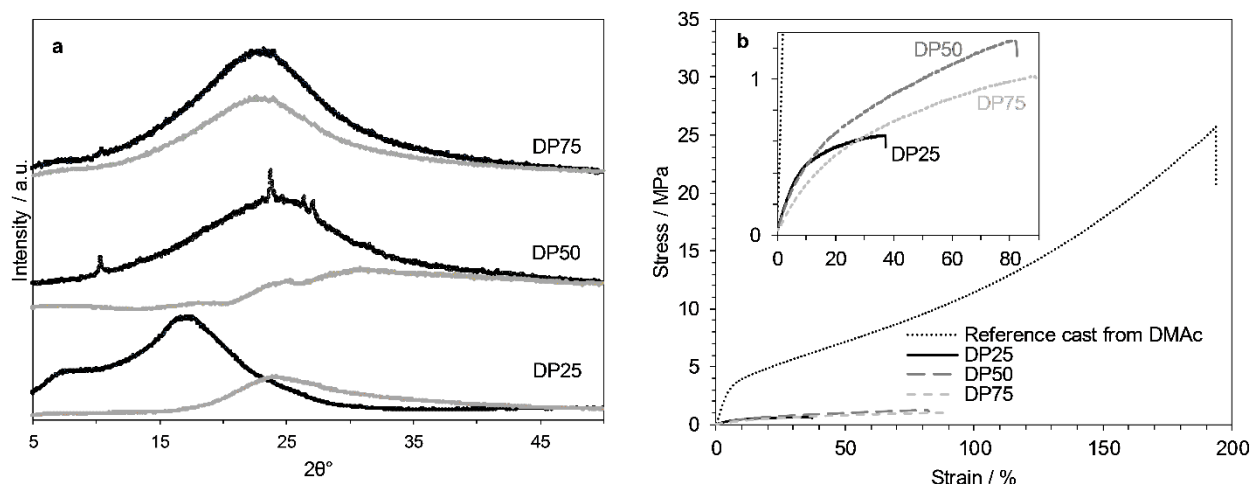


Figure 2. XRD of the membranes after the heat treatment (black) and after the subsequent equilibration in phosphoric acid (gray) (a). Representative stress-strain curves of the phosphoric acid equilibrated network membranes (b).

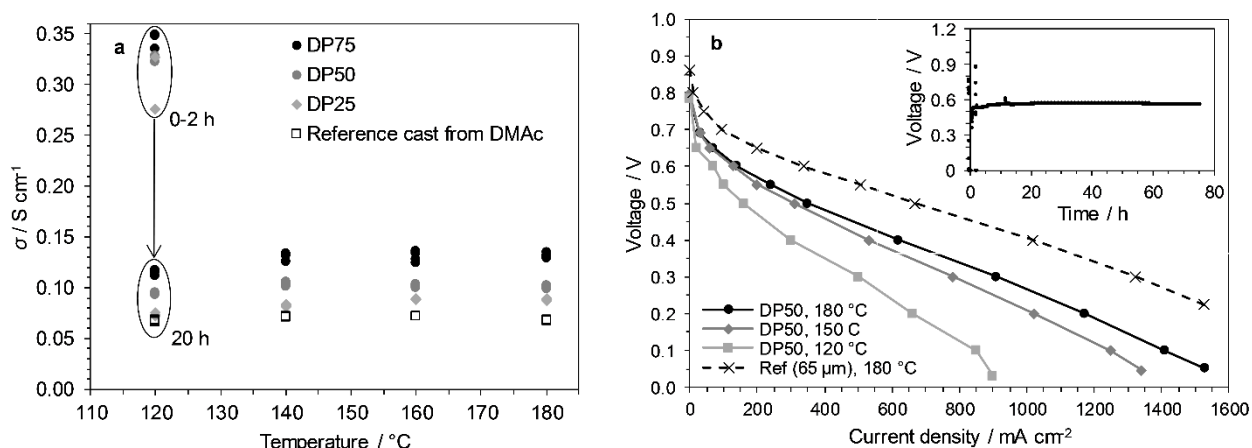


Figure 3. Proton conductivity of the *m*-PBI network membranes at 120-180 $^\circ C$ after an initial conditioning at 120 $^\circ C$ for 20 h (a) and fuel cell polarization curves of a MEA based on the DP50 membrane with a thickness of 300 μm (b). Fuel cell data for a cell based on linear *m*-PBI cast from DMAc with a doped thickness of 65 μm and 10-11 H_3PO_4 per polymer repeat unit are shown for comparison. The inset shows the cell voltage evolution at 200 $mA\ cm^{-2}$ for a cell based on the DP50 membrane operating at 160 $^\circ C$.

Table 1. Phosphoric acid contents of the membranes after equilibration in 85% PA. The corresponding acid doping levels are shown in parenthesis, normalized as H₃PO₄ per polymer repeat unit.

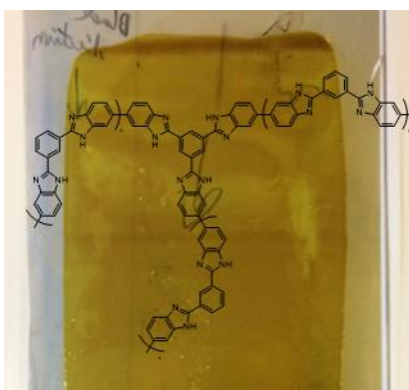
Membrane	Acid content / wt% (ADL ^a)		Elastic modulus / MPa
	Titration	Gravimetry	
DP25	94.6 (55.1)	94.0 (49.2)	6.6 ± 1.6
DP50	92.6 (39.3)	94.3 (52.0)	5.1 ± 0.3
DP75	91.3 (32.9)	92.0 (36.1)	3.1 ± 0.5

^a) Calculated based on a repeat unit molecular weight of 308 g mol⁻¹.

Covalent network polybenzimidazole membranes are obtained via a novel sequential direct casting procedure. Free-standing and mechanically robust gel electrolytes are obtained with phosphoric acid contents of up to 95 wt%. It supports high proton conductivity at temperatures well above 100 °C as ultimately manifested by low internal resistance in fuel cells operating at 180 °C.

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Gel electrolytes of covalent network polybenzimidazole and phosphoric acid by direct casting



Supporting Information

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Gel electrolytes of covalent network polybenzimidazole and phosphoric acid by direct casting

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Table S1. Experimental details for pre-polymer synthesis and membrane formation.

	DAB		IPA		BTA		PPA
	/ g	/ mmol	/ g	/ mmol	/ g	/ mmol	/ g
DP25	6.949	32.43	5.181	31.18	0.087	0.42	404
DP50	6.949	32.43	5.282	31.80	0.045	0.21	410
DP75	6.949	32.43	5.317	32.01	0.030	0.14	408

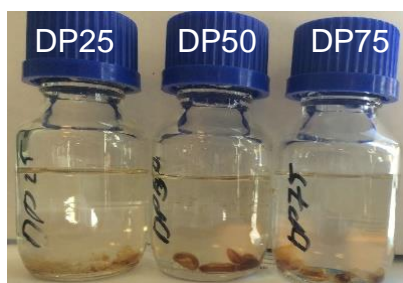


Figure S1. De-doped membranes after treatment in DMAc at 120 °C for 24 h.